

MyLab & Mastering |



An online tutorial system that helps students:

Study more efficiently

Improve results

Prepare for exam

Study in an interactive way

Practice for the course



MyLab&Mastering

Course ID

Assignment

Instructor Tools



To register you need:

- 1. Email address
- 2. Access code

e.g. XLFGH-LKJHG-XXXXX-HKGKS-AKR<mark>JG</mark>

To register students need:

- 1. Email address.
- 2. Access code
- 3. Course ID e.g. lecturer12345

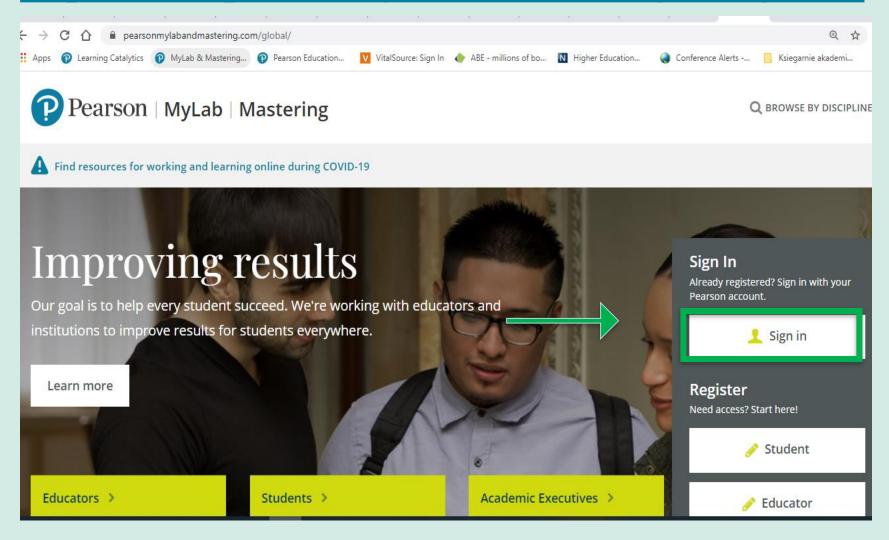






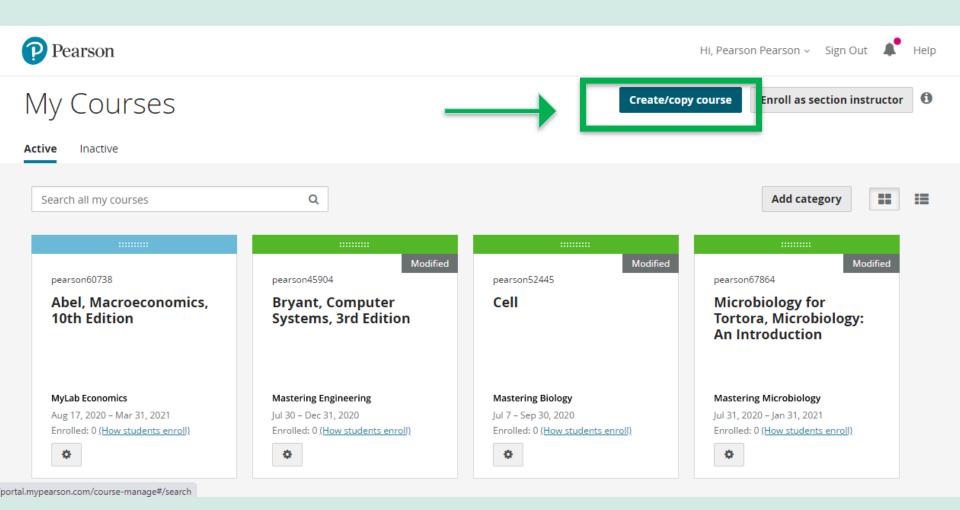
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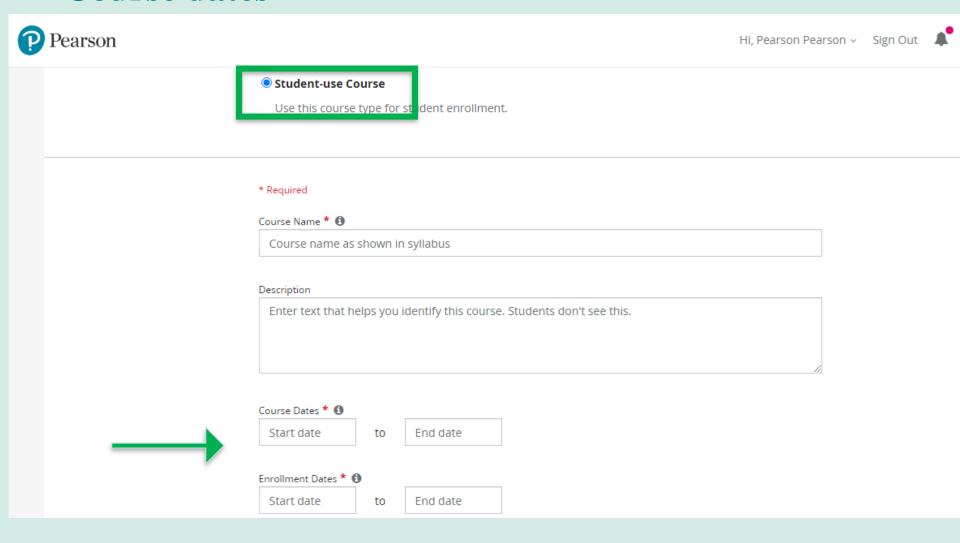


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Course dates



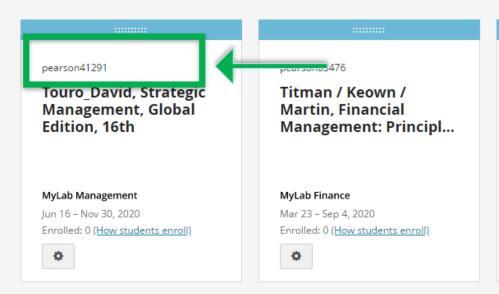


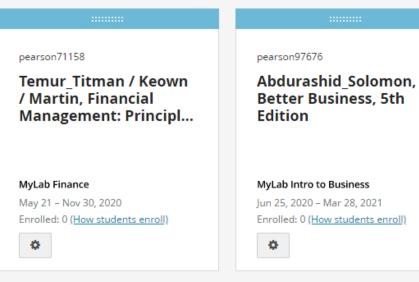
Know your course ID



Hi, Pearson Pearson V Sign Out







pearson39007 Robbins Organizational Behavior, 18

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5th Edition



Section instructors

Section instructors: Section instructors enroll with a Section Instructor access code at the Pearson MyLab and Mastering site.

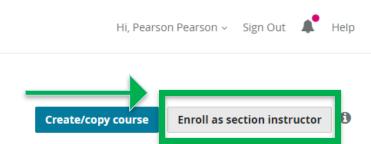
Section instructors are automatically granted almost all course privileges, so you may want to adjust these.)

Pearson See Section Instructor help for details.

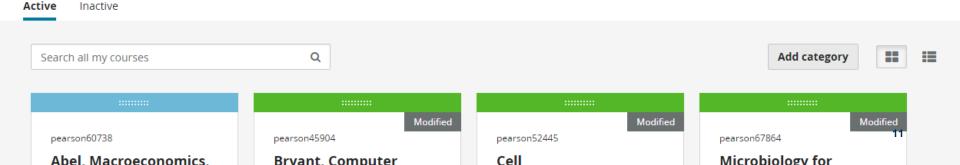
* Section instructor access code

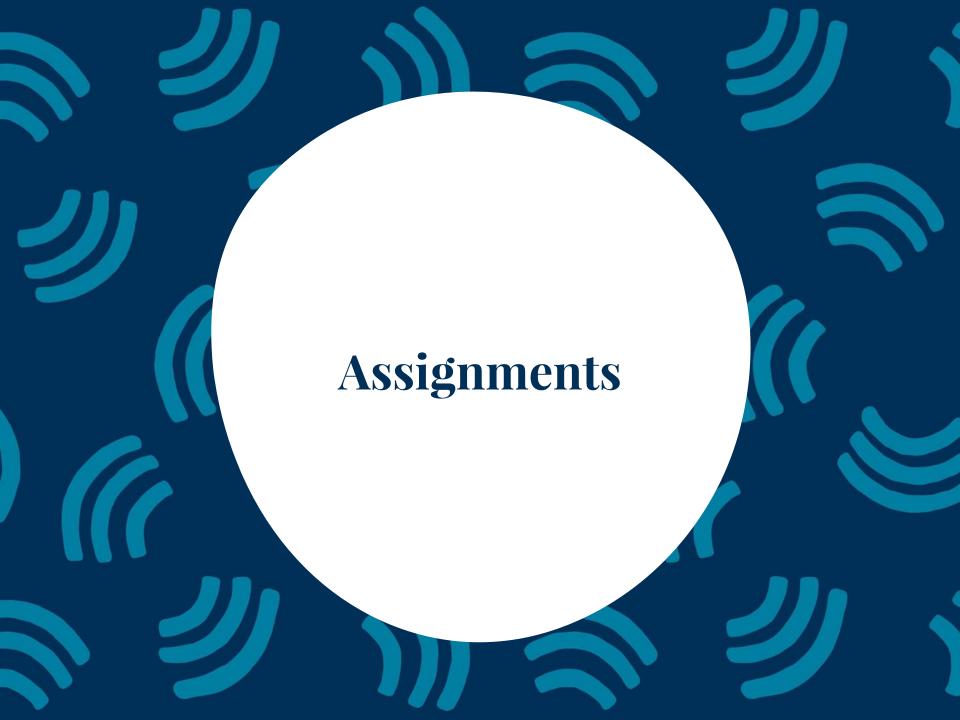
* Course ID

e.g. lecturer12345

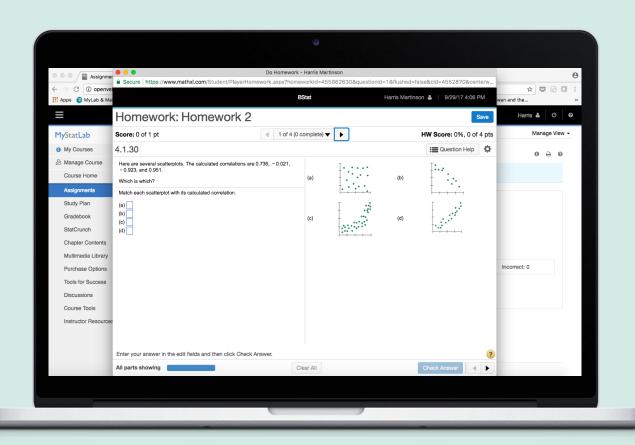


My Courses

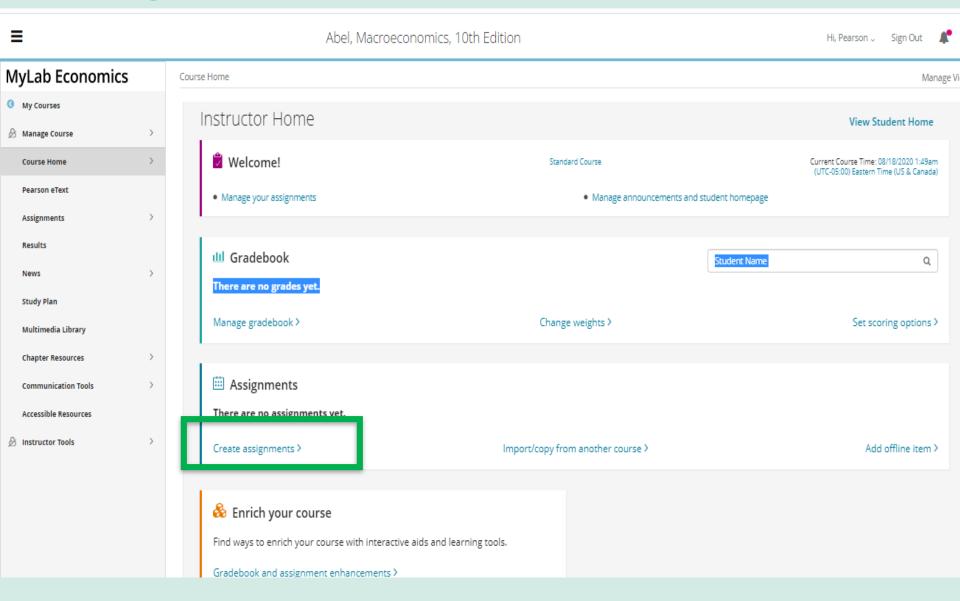




Assignments: Homework, Quizz, Test



Assignments



Assignments

These are assigned by you.

Students can open hints to help you answer the question.



Part A

If the reaction rate doubles when the temperature is increased to 35° C, what is the activation energy for this reaction in kJ/mol?

Express the activation energy in kilojoules per mole to two significant figures.



Incorrect; Try Again; 5 attempts remaining

Check that you have used a value of R and any conversion factors that cancel out undesired units so that the answer is in kJ/mol.

You may need to review Dimensional Analysis.

Hint 1. How to approach the problem

First, identify how the rate of reaction can be expressed in terms of the disappearance of reactant A. To do this, you can use the interactive activity. Click on the "Disappearance of reactants" button to see how the rate of reaction can be expressed in terms of the disappearance of reactant A for the given reaction.

Now, determine the change in the concentration of reactant A ($\Delta[A]$) during the given time interval t where t = 0 to 20 s by using the values of [A] given in the table. If $[A]_{20}$ is the concentration of reactant A at t = 20 s, and $[A]_0$ is the concentration of the reactant A at t = 0 s, then the change in concentration of reactant A is

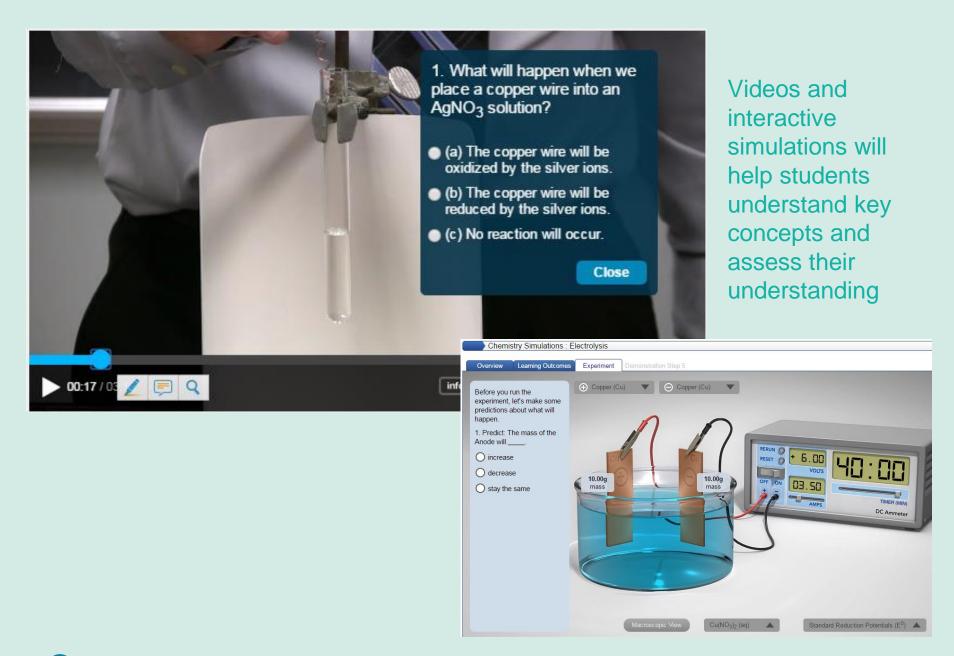
$$\Delta[A] = [A]_{20} - [A]_0$$

Once you know $\Delta[A]$, t, and the expression for the rate of reaction in terms of A, you can calculate the rate of reaction.



Students get feedback specific to the mistake they make so they can get it right next time

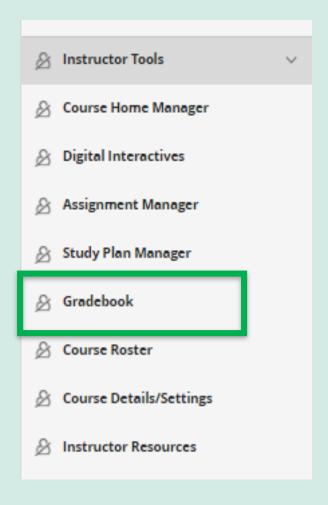








Instructor Tool



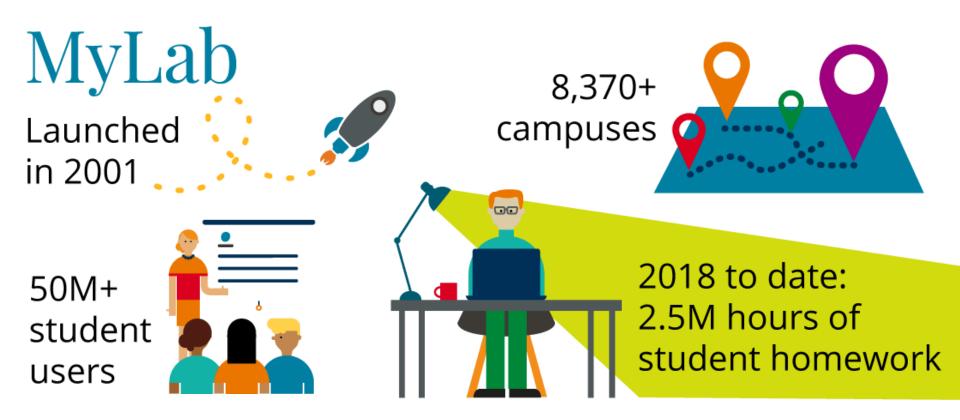


Gradebook

As an instructor, you would want to use the gradebook to:

- Access student results for a course or section
- Export grades and grade data into a spreadsheet
- Manage incomplete assignments, assignment weights, and other gradebook settings
- Understand how students are performing on specific assignment questions
- Conduct item analysis to understand more about each question in your assignments





Enabled just-in-time help for students learning math through problem-solving homework - now a fully personalized teaching and learning experience.

Available in Math, Business, professional training, and more.



Contact our helpdesk

Go to:

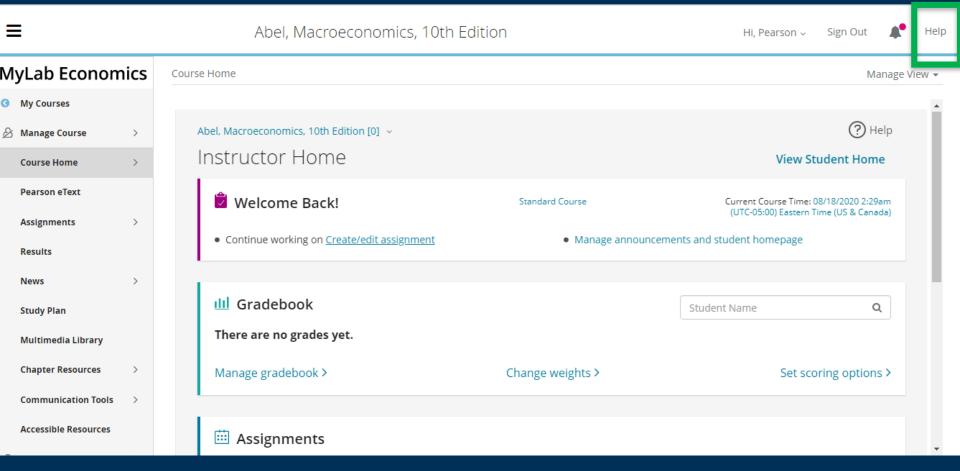
support.pearson.com/getsupport

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Help





ALWAYS LEARNING